AP Chemistry ts chapter 2 Atoms, Molecules, and Ions

Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

400 BC – Democritus, Greeks

Alchemy

Robert Boyle (1627-1691)

Stahl-phlogiston

Priestley

Lavoisier

Law of conservation of mass

Law of definite proportion

Dalton’s atomic theory

Atomic mass

Avogadro’s hypothesis

Berzelius, Selenium, and Silicon

J.J. Thomson

Millikan

Becquerel-radiation

Rutherford

Modern view of atomic structure

Nuclear symbol

Chemical bonds

Covalent bonds

Molecule

Chemical formula

Structural formula

Space filling model

Ball and stick model

Ionic bonding

Cation

Anion

Ionic solid

Polyatomic ion

Arrangement of the periodic table (atomic number, metals, nonmetals, groups or families, alkali metals, alkaline earth metals, halogens, noble gases, rows or periods)

Naming compounds

Binary ionic compounds – rules for naming

Writing formulas for binary ionic compounds

Binary ionic compounds type II

Ionic compounds with polyatomic ions – common polyatomic ions table 2.5

Binary covalent compounds type III

Rules for naming

Prefixes

Acids

Rules for naming

Review questions p. 68 1-10, 18, 20, 23,24,25,27,28,29,37,42,51,52,53,54,56,59,60,61, nomenclature #63-80. Marathon problem 112.